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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

Chapter 12.2: Stoichiometry of Reactions in Solution ...

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Key A In any stoichiometry problem, the first step is always to calculate the number of moles of each reactant present. In this case, we are Page 4/21. File Type PDF Chapter 12 Stoichiometry Practice Problems Answers given the mass of K 2

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water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample. 2. Saturn is about 1 429 000 km from the Sun.

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Chemistry: Supplemental Stoichiometry Problems. Directions: Solve each of the following problems. Assume excess of any reactant that isn't mentioned, unless otherwise . specified. Show your work, including proper units, to earn full credit. 1. Given the balanced equation, show what the following molar ratios (i.e., the conversion factors ...

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