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*Page 4/25*

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**Acids Bases**

CHAPTER 14 REVIEW .

Acids and Bases.

SHORT ANSWER

Answer the following questions in the space provided.

1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1:  
$$\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$$

b. Which stage of ionization usually produces more ions? Explain your answer.

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CHAPTER FOURTEEN  
ACIDS AND BASES

CHAPTER 14 ACIDS  
AND BASES 5

When  $\text{H}_3\text{PO}_4$  is added to water, the three acids that are present are  $\text{H}_3\text{PO}_4$ ,  $\text{H}_2\text{PO}_4^-$ , and  $\text{HPO}_4^{2-}$ .  $\text{H}_3\text{PO}_4$ , with the largest  $K_a$  value, is the strongest of these weak acids. The conjugate

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## **[DOC] Chapter 14 Review Acids And Bases Mixed**

CHAPTER 14 REVIEW  
Acids and Bases

SECTION 1 SHORT

ANSWER Answer the  
following questions in  
the space provided. 1.

Name the following  
compounds as acids:  
sulfuric acid a.  $\text{H}_2\text{SO}_4$   
sulfurous acid b.  $\text{H}_2\text{SO}_3$   
3 hydrosulfuric acid c.  
 $\text{H}_2\text{S}$  perchloric acid d.  
 $\text{HClO}_4$  hydrocyanic

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acid e. hydrogen cyanide 2. H<sub>2</sub>S Which (if any) of the acids mentioned in item 1 are binary acids? 3.

## **14 Acids and Bases**

-Acids are defined as any substance that can accept an electron pair to form a covalent bond (Substances do NOT have to contain Hydrogen) -Bases are defined as any substance that donates an electron pair



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CHAPTER 14 REVIEW.

Acids and Bases.

SHORT ANSWER

Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases:  $\text{HSO}_3^-$ . 3 -. a. What is the conjugate base of  $\text{H}_2\text{SO}_4$ .

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Acids, Bases and Salts  
1. Know the meanings  
of and be able to apply  
the following terms:  
Bronsted-Lowry acid

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strong acid

concentrated chemical

...

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Textbook Authors:  
Zumdahl, Steven S.;  
Zumdahl, Susan A. ,  
ISBN-10: 1133611095,  
ISBN-13:  
978-1-13361-109-7,  
Publisher: Cengage  
Learning

## **Chapter 14 - Acids and Bases - Review Questions - Page 699: 2**

When an acid and a  
base are mixed, the  
 $H^+$  from the acid  
combines with the  $OH^-$

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from the base to form  $H_2O$ . Acid-Base Titration In a titration, a substance in a solution of KNOWN concentration is reacted with another substance in a solution of UNKNOWN concentration.

## **Acids and Bases** **Chapter 14** **Flashcards | Quizlet**

CHAPTER 14 ACIDS  
AND BASES 5 When  
 $H_3PO_4$  is added to

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water, the three acids that are present are  $\text{H}_3\text{PO}_4$ ,  $\text{H}_2\text{PO}_4^-$ , and  $\text{HPO}_4^{2-}$ .  $\text{H}_3\text{PO}_4$ , with the largest  $K_a$  value, is the strongest of these weak acids. The conjugate bases of the three acids are  $\text{H}_2\text{PO}_4^-$ ,  $\text{HPO}_4^{2-}$

## **CHAPTER FOURTEEN ACIDS AND BASES**

Proton transfer reactions favor the production of [stronger/weaker]

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CHAPTER FOURTEEN  
ACIDS AND BASES For  
Review 1. ... CHAPTER  
14 ACIDS AND BASES 5  
When  $\text{H}_3\text{PO}_4$  is added  
to water, the three  
acids that are present  
are  $\text{H}_3\text{PO}_4$ ,  $\text{H}_2\text{PO}_4^-$ ,  
and  $\text{HPO}_4^{2-}$ .  $\text{H}_3\text{PO}_4$ ,  
with the largest  $K_a$   
value, is the strongest  
of these weak acids.



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The conjugate ...

**Modern Chemistry  
Chapter 14 Review  
Answers Acids And  
Bases**

Acids and Bases:

Chapter 14 & 15 . HW:

•Read Ch 14: •Fill in as  
much of the acid base  
table as you can, as  
you read . Acid base  
conductivity and  
reactivity Conductivity,  
Reac'vity,, ... 3.Work  
on ch 14-15 review  
(due Wednesday, test

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Friday!) • Solutions in  
which [H

## **Acids and Bases: Chapter 14 & 15**

Chapter 14: Acids and  
Bases. Ch 14 Page 1.

Arrhenius Base -a  
hydroxide containing  
compound that

dissociates to produce  
OH-when dissolved in  
water  $\text{NaOH}(\text{aq})$

$\rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$

Bases that contain OH-  
are ionic compounds.

Ionic substances

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dissociate in water.

**Chapter 14: Acids  
and Bases**

pH of a weak base: pH  
of .2 M of  $\text{NH}_3$  (weak  
base). Conjugate acids  
and bases :

Introduction to  
conjugate acids and  
bases  $\text{pK}_a$  and  $\text{pK}_b$   
relationship : The  $\text{pK}_a$   
and  $\text{pK}_b$  relationship  
between conjugate  
acids and bases (both  
of which are weak).

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Acids Bases

**Chapter Fourteen  
[Acids and Bases] -  
Wattsburg**

Modern Chemistry 121

Acids and Bases

CHAPTER 14 REVIEW

Acids and Bases

SECTION 3 SHORT

ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: \_\_\_\_\_ a. What is the conjugate base of

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H<sub>2</sub>SO<sub>3</sub>? \_\_\_\_\_ b.

**CHAPTER 14 REVIEW**  
**Acids and Bases**

equivalent amounts of acids and bases react, the three properties just described disappear because the acid is “neutralized.” The reaction products are water and an ionic compound called a salt. 5. Acids conduct electric current. ... 3.

CHAPTER 14 4. Bases

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**CHAPTER 14 Acids  
and Bases**

Chapter 14 Review:  
Acids and Bases.

Section 14.1 Bronsted  
Acids and Bases. Acids  
- molecules that can  
lose  $H^+$  (proton donors)  
making  $H_3O^+$  in water  
(acids lose  $H^+$ ) Bases -  
molecules than can  
gain  $H^+$  (proton  
acceptors) and/or  
make  $OH^-$  in water  
(bases gain  $H^+$  / make

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OH-) Conjugate  
acid/base pairs - differ  
by one H+.

## **Chapter 15 Review Acid Base Equilibria**

Chapter 14 Acids Bases  
Answers Chapter 14  
Acids Bases Answers

As recognized,  
adventure as capably  
as experience nearly  
lesson, amusement, as  
competently as  
settlement can be  
gotten by just checking  
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