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## **Chapter 19 Acids Bases And**

An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous

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solution. Bronsted-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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conductors of electricity-can turn litmus paper a different color (red)-reacts with certain metals-contains more hydrogen ions than hydroxide ions-reacts with carbonates. Bases-taste bitter-conductors of electricity

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produce.... what did geologist find  
interesting abo.... -taste sour...

-conductors of electricity... -can turn  
litmus paper.... -taste bitter...

-conductors of electricity... -feels

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slippery... -can....

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The Arrhenius definition states that acids are compounds that contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen ( $H^+$  ions). Bases are compounds that ionize to yield hydroxide ions ( $OH^-$ ) in aqueous solutions.

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to compare, contrast, and evaluate the nature of acids and bases in terms of percent dissociation, hydronium ion concentration, and pH.

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Chapter 19 Acids, Bases, and Salts Acids and Bases: Basics These are really just a specific type of chemical compound.

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Answer C, a conjugate base is what is  
left of an ion after it has donated a  
proton in solution.

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## Workbook Answers SECTION 19.1 ACID-BASE

THEORIES(pages 587-593) This section compares and contrasts acids and bases as defined by the theories of Arrhenius, Brønsted-Lowry, and Lewis. It also identifies conjugate acid-base pairs in acid-base reactions.

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