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David W. Ball is Professor of Chemistry at Cleveland State University. His research interests include computational chemistry of new high energy materials, matrix isolation spectroscopy, and various topics in chemical education.

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First, calculate the energy needed to warm the water by 1.00°C:  $q = mcT = (1.00105 \text{ g})(4.18 \text{ J/gK})(1.00 \text{ K}) = 4.18105 \text{ J}$  Now, determine how many drops of a 20.0 kg weight falling 2.00 meters in gravity will yield that much energy. The amount of energy in one drop is  $mgh = (20.0 \text{ kg})(9.81 \text{ m/s}^2)(2.00 \text{ m}) = 392.4 \text{ J}$ .

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